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gas constant and the ratio  
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and mass flow rate  $\text{kg/s}$ .  
Assuming that the

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temperature of the oxygen in the bottle does not change with time, determine the time it takes to reduce the pressure to one half of its initial value.  $V=0.1 \text{ m}^3$   $p_1=10 \text{ MPa}$   $T_1=293\text{K}$   $=T_2$   $p_2=5 \text{ MPa}$   $\text{kg}$   $\text{K}$   $\text{J}$



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