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~~the type of mixture on the left to its representative particle diameter on the right. c solutions (a) larger than 1000 nm a~~
~~suspensions (b) 1 nm to 1000 nm b colloids (c) smaller than 1 nm 2. Identify the solvent in each of the following examples:~~

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~~ANSWER 1. c 2. a 3. b 2. a. alcohol b. water c. the gels 3. The mixture is a colloid. The properties are consistent with those~~
~~reported in Table 3 on page 404 of the text. The particle size is small, but not too small, and the mixture~~

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Modern Chemistry Chapter 12 Mixed CHAPTER 12 REVIEW Solutions MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Solid CaCl₂ does not conduct electricity. Explain why it is considered to be an electrolyte. CaCl₂ is an ionic solid. In the crystal form, its ions are locked in position. Dissolving mc06se cFMsr i-vi

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